
PHY2505S

Atmospheric Radiative Transfer and Remote Sounding

Lecture 8

- Molecular Absorption and Emission
- Vibration-Rotation Spectra

Greenhouse Gases

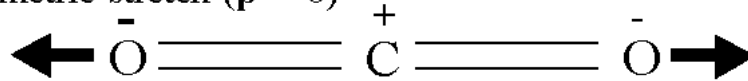
- **Water vapour (H₂O)**
 - most common greenhouse gas
 - increases as surface temperature rises
- **Carbon dioxide (CO₂)**
 - released by plant and animal life, decay, and burning of fuels
 - removed by plant photosynthesis and absorption by the oceans
- **Methane (CH₄)**
 - not as common in volume as H₂O or CO₂
 - very effective at trapping heat - powerful greenhouse gas
 - wetlands, rice paddies, animal digestion, fossil fuel extraction, decaying garbage
- **Nitrous oxide (N₂O)**
 - soils and the oceans, some from burning fossil fuels and fertilizer use
- **Ozone (O₃)**
 - most ground level ozone is from chemical reactions involving pollutants
- **Halocarbons**
 - anthropogenic chemicals containing bromine, chlorine, or fluorine, and carbon
 - extremely powerful greenhouse gases

Greenhouse Gases (GHGs)

Greenhouse gases = gases with vib-rot absorption features at $\sim 5\text{-}50\ \mu\text{m}$

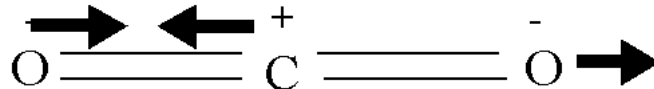
VIBRATIONAL MODES OF CO_2

1. Symmetric stretch ($p = 0$)



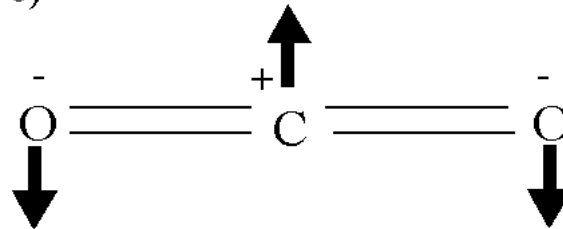
**IR inactive – no change
in the dipole moment**

2. Asymmetric stretch ($p \neq 0$)



IR active

3. Bend ($p \neq 0$)

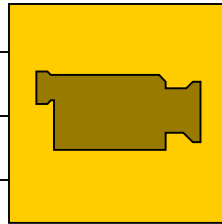


IR active

- Major greenhouse gases: H_2O , CO_2 , CH_4 , O_3 , N_2O , CFCs, ...
- Not greenhouse gases: N_2 , O_2 , Ar, ...

Adapted from D. Jacob

Molecular Absorption and Emission



Rotational Energy Levels

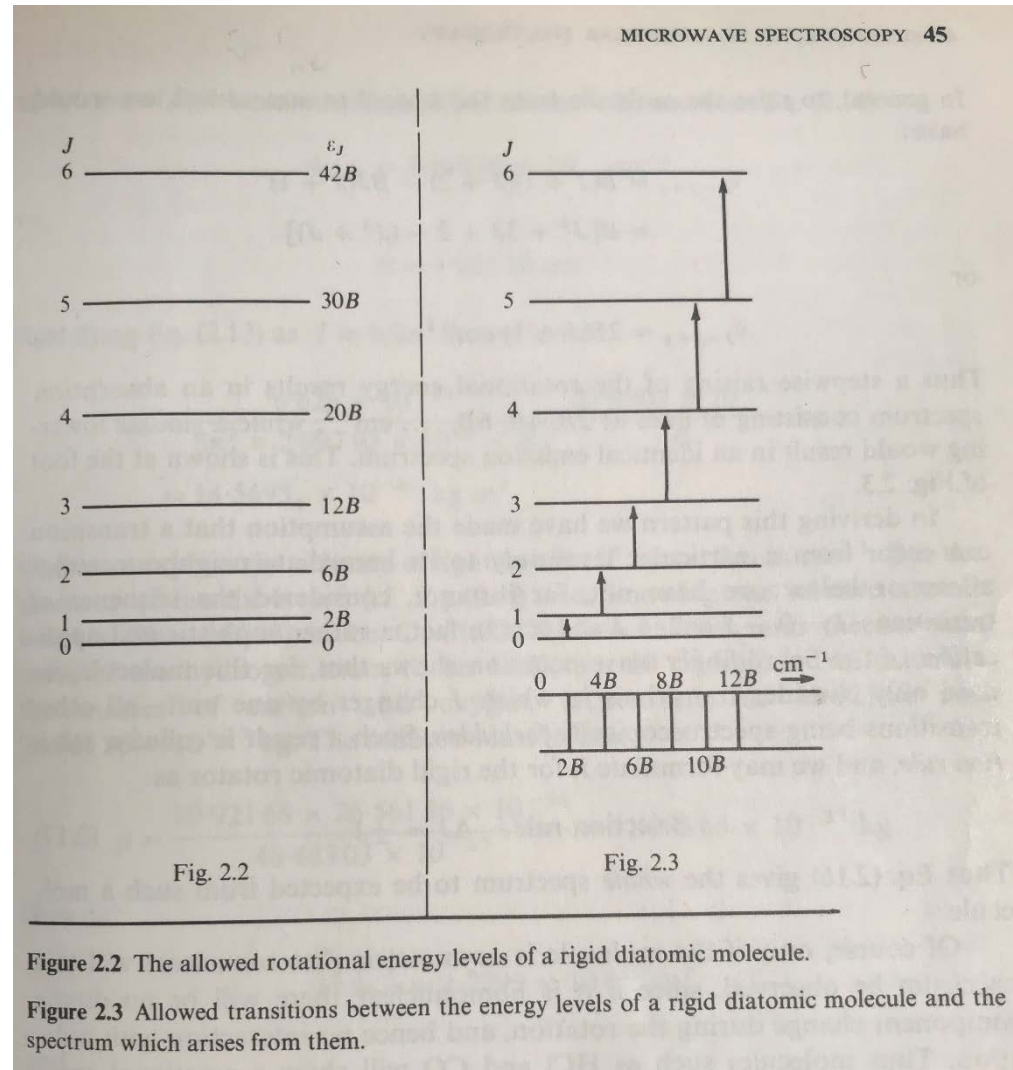
Figures are from C.N. Banwell,
Fundamentals of Molecular
Spectroscopy, 3rd Edition, 1983.

These diagrams are relevant to
material on pages 2-3 of the notes.

Figure 2.2 shows the energy levels:
 $E = BJ(J+1)$ in cm^{-1}
 $= 0, 2B, 6B, 12B, 20B, \text{etc.}$

Figure 2.3 shows the locations of
spectral lines based on energy
transitions:

$\Delta E = 2B(J+1)$ in cm^{-1}
 $= 2B, 4B, 6B, 8B, \text{etc.}$
so they are equally spaced



Vibration-Rotation Spectra

